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For instance,  $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$

denotes that there are four atoms of hydrogen and 2 atoms of oxygen on both sides of the equation. The amount of reactants must be equal to

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the amount of products. When students get big chemical equations in a balancing equation worksheet, they often find it to be very difficult.

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Balancing Chemical Equations.

Balance and classify five types of chemical reactions: synthesis, decomposition, single replacement, double replacement, and combustion.

While balancing the reactions, the

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number of atoms on each side is presented as visual, histogram, and numerical data.

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Balance: To balance a chemical equation, you are not allowed to

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Change the chemical formulas of the substances involved in the reaction. You are allowed to change the number of molecules of each...

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## Answer Key

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equations is important when trying to understand nuclear reactions. All equations need to be balanced to conform to two conservation laws: the mass number is conserved, and the



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Electrical charge is conserved.

Success Criteria Use the conservation laws to find an unknown in a nuclear reaction equation. Write a balanced nuclear equation for a natural transmutation. Prerequisites

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## Home Answer Key

To balance a chemical equation, enter an equation of a chemical reaction and press the Balance button. The balanced equation will appear above. Use uppercase for the first character in the element and lowercase for the second character. Examples: Fe, Au,

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Co, Br, C, O, N, F. Ionic charges are not yet supported and will be ignored.

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Balanced equations are those whose coefficients result in equal numbers of atoms for each element in the

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reactants and products. Chemical reactions in aqueous solution that involve ionic reactants or products may be represented more realistically by complete ionic equations and, more succinctly, by net ionic equations.

## **4.1: Writing and Balancing**

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Stoichiometry. Balancing chemical equation with substitution.

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