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Many companies, including pharmaceutical companies, keep track of the concentration of a solution by measuring its molarity — a ratio of the number of solute particles to the volume of the solution. In this activity you will learn about molarity and how to represent concentration quantitatively. Model 1 — Lemonade Mixtures*

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of the new solution? (Answer: 9-801 HN03) = ML (0.8M -N
3-30.00 ml, of water is added to 20.00 ml, of 0.0500 M sodium carbonate. What is the molarity of sodium carbonate in the new solution? (Answer: 0.0200 M) (20.00 m C (350. 00 casF) 4. What volume of 0.100 M NaCl is needed to make 5.00 ml, of 0.200 M solution? (Answer: 0.01 (0.100 M) V)

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